**Plant Operator in Training Mathematics and Science Study Guide**

An Operator-in-training is responsible for learning to operate all areas of a water or wastewater treatment plant.

**MATHEMATICS**

We suggest you are familiar with multiplication, division, fractions, average and median, and basic algebraic formulas, in addition to the items listed below in order to be properly prepared for the math section of this exam.

Note the abbreviations used:

A – Area l – length V – Volume

b – base r -radius

h – height w – width

**Area Formulas -** Circle A= *π* r2 Rectangle A= lw Triangle A= ½ bh

**Volume and Surface Area Formulas**

 **Volume** **Surface Area**

Cube V= A3 S.A. = 2wl + 2lh + 2hw

Rectangular Prism V= lwh S.A. = 2(lw + lh + wh)

Cylinder V= *π* r2h S.A. = 2πrh + 2πr2

Example: A tank 4 feet wide, 7 feet long, and 10 feet high is filled with water to the exact top. The volume in cubic feet is

1. 70
2. 120
3. 280 Correct answer
4. 360

In algebraic equations, the Order of Operations is, in this order; **P**arenthesis, **E**xponents, **M**ultiplication, **D**ivision, **A**ddition, and **S**ubtraction. An easy way to remember this is the Acronym **PEMDAS.**

Example: In the equation (CxD)+(AxB)-F, the first mathematical operation performed is

1. subtraction
2. multiplication Correct answer
3. addition
4. division

**Types of slope**



**Science**

We suggest you are familiar with atoms, their charges, nuclear change, chemical formulas, plant cells and their components, in addition to the items listed below in order to be properly prepared for the science section of this exam.

**Atoms**

Components of an atom - [electrons](http://www.chem4kids.com/files/atom_electron.html), protons, and [neutrons](http://www.chem4kids.com/files/atom_neutron.html). Electrons are the smallest of the three particles that make up atoms. Electrons are found in shells or [orbitals](http://www.chem4kids.com/files/atom_orbital.html) that surround the nucleus of an atom. Protons and neutrons are found in the **nucleus**. They group together in the center of the atom.

Atomic Number - the number of protons in the nucleus of an atom, which determines the chemical properties of an element and its place in the periodic table.

Atomic Charge - Electrical **charge** density due to gain or loss of one or more electrons.

Types of reactions

Chemical Reaction - a process in which one or more substances, the reactants, are converted to one or more different substances, the products.

Physical Reaction - occurs when molecules undergo a molecular rearrangement to produce a physical change. The molecules are not chemically altered.

Nuclear reaction - a change in the identity or characteristics of an atomic nucleus that results when it is bombarded with an energetic particle, as in fission, fusion, or radioactive decay.

**States of matter**

|  |
| --- |
| **Some Characteristics of Gases, Liquids and Solids and the Microscopic Explanation for the Behavior** |
| **gas** | **liquid** | **solid** |
| assumes the shape and volume of its containerparticles can move past one another | assumes the shape of the part of the container which it occupiesparticles can move/slide past one another | retains a fixed volume and shaperigid - particles locked into place |
| compressiblelots of free space between particles | not easily compressiblelittle free space between particles | not easily compressiblelittle free space between particles |
| flows easilyparticles can move past one another | flows easilyparticles can move/slide past one another | does not flow easilyrigid - particles cannot move/slide past one another |

**Photosynthesis** - the process by which green plants and some other organisms use sunlight to synthesize foods from carbon dioxide and water. Photosynthesis in plants generally involves the green pigment chlorophyll and generates oxygen as a byproduct.

**MATHEMATICS FORMULAS**

**Area Formulas**

Parallelogram A= bh Trapezoid A= ½ h(b1+b2) Circle A= *π* r2

Rectangle A= lw Triangle A= ½ bh

**Circumference of a circle**

C= *π*d or C= 2*π*r (note: *π= 3.14 or 22/7)*

**Volume and Surface Area Formulas**

 **Volume** **Surface Area**

Cube V= A3 S.A. = 2wl + 2lh + 2hw

Sphere V= 4/3 *π* r3 S.A. = 4πr2

Rectangular Prism V= lwh S.A. = 2(lw + lh + wh)

Pyramid V= ⅓ bh S.A. = ½ pl (p = perimeter, l = slant)

Cone V= ⅓ *π* r2h S.A. = πr(r + √h2 + r2)

Cylinder V= *π* r2h S.A. = 2πrh + 2πr2

**U.S. Customary Measure**

|  |  |  |
| --- | --- | --- |
| **Length** | **Liquid** | **Weight** |
| 1 foot = 12 inches | 1 cup = 8 fluid ounces | 1 pound = 16 ounces |
| 1 yard = 3 feet | 1 pint = 2 cups | 1 ton = 2000 pounds |
| 1 mile = 5280 feet | 1 quart = 2 pints |  |
|  | 1 gallon = 4 quarts |  |

**Metric Measurement**

|  |  |  |
| --- | --- | --- |
| **Length** | **Liquid** | **Weight** |
| 1 centimeter = 10 millimeters | 1 centiliter = 10 milliliters | 1 centigram = 10 milligrams |
| 1 meter = 100 centimeters | 1 liter = 100 centiliters | 1 gram = 100 centigrams |
| 1 kilometer = 1000 meters | 1 kiloliters = =1000 liters | 1 kilogram = 1000 grams |

**Selected Conversions of U.S. Customary Units to Metric Units**

|  |  |  |
| --- | --- | --- |
| **Length** | **Liquid** | **Weight** |
| 1 inch = 2.54 centimeters | 1 quart = .95 liters | 1 ounce = 28.35 grams |
| 1 mile = 1.61 kilometers | 1 gallon = 3.79 liters | 1 ton = 907.18 kilograms |

**Temperature Conversion: Centigrade Fahrenheit**

 0C = 5/9 (0F-32) 0F = 9/5 (0C+32)